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704 CHAPTER 22 Identify the product that balances the following nuclear reaction: ${}_{84}^{212}\text{Po} \rightarrow {}_{82}^{208}\text{Pb} + {}_2^4\text{He}$ . The total mass number and atomic number must be equal on both sides of the equation.  ${}_{84}^{212}\text{Po} \rightarrow {}_{82}^{208}\text{Pb} + {}_2^4\text{He}$  + mass number:  $212 = 208 + 4$  atomic number:  $84 = 82 + 2$

2. The nuclide has a mass number of 208 and an atomic number of 82,  ${}_{82}^{208}\text{Pb}$ . 3.

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Define and relate the terms radioactive decay and nuclear radiation  
Describe the different types of radioactive decay and their effects on  
the nucleus Define the term half-life, and explain how it relates to  
the stability of a nucleus Define and relate the terms decay series,  
parent nuclide ...

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Use dimensional analysis and the given densities to make the

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following conversions. a. 14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/cm<sup>3</sup>.  $(14.8 \text{ g}) \left( \frac{1 \text{ cm}^3}{2.34 \text{ g}} \right) = 6.32 \text{ cm}^3$  b. 2.8 L of argon (Ar) to grams of argon.

## ~~Chapter 4 Problem Solving in Chemistry Answer Key~~

1) No, it is not a proper answer; you do not know whether the professor meant homework problem number 20 or 20 homework problems. Twenty is a quantity and does not make sense without a unit. 3) Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

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